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Chapter 13 - Gases - An Introduction to Chemistry

Chapter 13 Study Guide: Gases. ... 13. What is the pressure of dry hydrogen gas collected over water if the temperature is 23°C and the pressure is 842 torr? 14. 2 C 2 H 6 + 7 O 2--> 4 CO 2 + 6 H 2 O. How many milliliters of ethane (C 2 H 6) at STP are required ...

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194 Study Guide for An Introduction to Chemistry Section Goals and Introductions Section 13.1 Gases and Their Properties Goals • To describe the particle nature of both real and ideal gases. • To describe the properties of gases that can be used to explain their characteristics: volume, number of particles, temperature, and pressure.

Chapter 13 - Gases

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Section 13.6 Exercise 27.4 L of oxygen gas at 25.0°C and 1.30 atm, and 8.50 L of helium gas at 25.0°C and 2.00 atm were pumped into a tank with a volume of 5.81 L at 25°C. • Calculate the new partial pressure of oxygen. 6.13 atm • Calculate the new partial pressure of helium. 2.93 atm • Calculate the new total pressure of both gases. 9 ...

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Study Guide Chapter Assessment Standardized Test Practice Gases Resources CHAPTER 13 . ... 13.1 The Gas Laws Study Guide . Key Concepts • Gay-Lussac’s law states that the pressure of a fixed amount of gas is directly proportional to its kelvin temperature at constant volume.

Chemistry: Matter and Change

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A gas is a state of matter with no defined shape or volume. Gases have their own unique behavior depending on a variety of variables, such as temperature, pressure, and volume. While each gas is different, all gases act in a similar matter. This study guide highlights the concepts and laws dealing with the chemistry of gases.

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Chapter 13 Study Guide Gases Chapter 13 – Gases 195 Exercise 13.3 – Equation Stoichiometry: Iron is combined with carbon in a series of reactions to form pig iron, which is about 4.3% carbon. 2C O2 2CO Fe2O3 3CO 2Fe 3CO2 2CO C (in iron) CO2 Pig iron is easier to shape than pure iron, and the presence of carbon lowers its melting point

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CHAPTER SOLUTIONS MANUAL GasesGases Solutions Manual Chemistry: Matter and Change • Chapter 13 253 Section 13.1 The Gas Laws pages 442–451 Practice Problems page 443 Assume that the temperature and the amount of gas are constant in the following problems. 1. The volume of a gas at 99.0 kPa is 300.0 mL. If

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Chapter 13 Study Guide Answers 1. Describe the assumptions/postulates of the kinetic-molecular theory of gases: a. Gases are composed of tiny particles in constant rapid, (“random” or “straight–lined”?) motion. b. Gases are separated by relatively huge distances. The volume of the particles is essentially zero. c.

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Chapter 13: Standard Review Worksheet 1. While the barometer is used to measure atmospheric pressure, a device called a mercury manometer is used to measure the pressure of samples of gas in the laboratory.A manometer consists basically of a U-shaped tube filled with mercury, with one arm of the

Chapter 13: Standard Review Worksheet

Study Guide Chemistry: Matter and Change • Chapter 13 19 Section 13.2 The Combined Gas Law and Avogadro’s Principle In your textbook, read about the combined gas law. Fill in the following table. State what gas law is derived from the combined gas law when the variable listed in the first column stays constant and the variables in the

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